

Gaining Apex Coaching Centre

(Where Toppers make..... Toppers)

Compiled By: Dapinderjeet Singh

(10+1 CHEMISTRY)

Periodic Table (Assignment -I)

- 1) Why radius of cation is smaller than that of parent atom while that of anion is larger than that of parent atom
- 2) Halogens have maximum negative electron gain enthalpy Why
- 3) What are the limitation of Mendeleev's periodic table?
- 4) What is the cause of periodicity in periodic table
- 5) The atomic number of an element is 24 at which place it is placed in the periodic table
- 6) Write the general Electronic configuration of d and f block elements?
- 7) What is Ionization Energy? On what factors does it depend.
- 8) The negative Electron gain Enthalpy of Fluorine is less than that Chlorine. Explain
- 9) Noble gases have almost zero or positive or negative electron gain enthalpy
- 10) What is diagonal relationship and explain its cause
- 11) Alkali metal do not forms dispositive ions?
- 12) Would you expect the second electron gain enthalpy of oxygen as positive, less negative or more negative that first? Justify your answer.
- 10) First Ionization energy of Mg is more as compared to that of Na while its second ionization energy is less?
- 11) Would you except the Ionization enthalpies of two isotopes of the same element to be same or different? Justify your answer
- 12) Why the I.E.₂ of the element is more than that of I.E.₁
- 13) Oxygen has lower $\Delta_f H_1$ than N and F . Why?
- 14) Out of O and F which has higher value of negative electron gain enthalpy
- 15) Out of Be and B which has lesser Electron gain enthalpy and why?
- 16) Chlorine can be converted easily to chloride ion more easily as compared to fluoride ion from fluorine
- 17) Ne and Na⁺ are isoelectronic species. Do they have same ionization enthalpy also
- 18) Define Electron gain enthalpy and on what factors does it depends
- 19) Third period has eight elements not 18. Explain
- 20) Arrange the following elements in increasing order of their metallic character B,Al,Mg,K
- 21) Arrange the following elements in increasing order of their non metallic character B, C,Si,N,F
- 22) Which of the following would have larger size i) K or K⁺ ii) O²⁻ or F⁻ iii) Li⁺ or Na⁺ iv) Na⁺ or Mg²⁺ v) P or As
- 23) Consider the following species N³⁻, O²⁻,F⁻, Na⁺,Mg²⁺,Al³⁺ What is common in them and Arrange them in order of increasing radii
- 24) Arrange them LiCl, NaCl, KCl, RbCl in order of increasing melting point