Gaining Apex Coaching Centre

(Where Toppers make...... Toppers)



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	a) 3.01×10^{24}	b) 4.06×10^{24}	c) 2.01×10^{24}	d) 3.02×10^{24}
8.	In the following reaction, $MnO_2 + 4HCL \rightarrow MnCl_2 + 2H_2O + Cl_2$			
	a) 25%	b) 50%	c) 100%	d) 75%
9.	The normality of 1% (va) 0.02	vt./vol.)H ₂ SO ₄ is nearly: b)0.2	c) 0.1	d) 1
10.	The mass of 1 mole of e a) 9.1×10^{-28} g	electrons is b) 1.008 mg	c) 0.55 mg	d) 9.1 × 10^{-27} g
11.	74.4 g of a metallic chlo a) 19.5	oride contains 35.5 g of c b) 35.5	chlorine. The equivalent c) 39.0	weight of the metal is: d) 78.0
12.	Equivalent weight of an acid a) Depends on the reaction involved b) Depends upon the number of oxygen atoms present c) Is always constant d) None of the above			
13.	Which of the following a) Gasoline	is not a mixture? b)Distilled alcohol	c) LPG	d) lodized table salt
14.	The equivalent weight of a divalent metal is 31.82. The weight of single atom is: a) $32.77 \times 6.02 \times 10^{23}$ b) $63.64 \times 6.02 \times 10^{23}$ c) 63.64 d) $63.64/6.02 \times 10^{23}$			
15.	Number of mole of 1 m a) 44.6	³ gas at NTP are: b) 40.6	c) 42.6	d) 48.6
16. The per cent loss in weight after heating a pure sample of potass122.5) will be:				lorate (mol. wt. =
	a) 12.25	b)24.50	c) 39.18	d)49.0
17.	The number of milli eq a) 0.1	uivalent contained in 0.5 b) 100	5 litre of 0.2 <i>N</i> solution is c) 0.01	s: d) 1.0

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- 18. Out of 1.0 g dioxygen, 1.0 g (atomic) oxygen and 1.0 g ozone, the maximum number of molecules are contained in
 a) 1.0 g of atomic oxygen
 b) 1.0 g of ozone
 c) 1.0 g of oxygen gas
 d) All contain same number of atoms
- 19. A sample of AIF_3 contains 3.0×10^{24} F ions. The number of formula units of this sample area) 9.0×10^{24} b) 3.0×10^{24} c) 0.75×10^{24} d) 1.0×10^{24}
- 20. One mole of CO_2 contains a) 3 g atoms of CO_2 c) 6.02×10^{23} atoms of 0

b) 18.1 \times 10 23 molecules of CO $_2$ d) 6.02 \times 10 23 atoms of C